

CH301 Final Exam equations, constants, conversions

Constants and Conversions:

$$1 \text{ atm} = 1.013 \times 10^5 \text{ Pa} = 760 \text{ torr} = 760$$

mmHg

$$1000 \text{ g} = 1 \text{ kg}$$

$$R = 0.082 \text{ l atm/K mol}$$

$$R = 1.987 \text{ cal/mol K}$$

$$R = 8.314 \text{ J/mol K}$$

$$k = 1.38 \times 10^{-23} \text{ J/K}$$

$$N = 6.022 \times 10^{23}$$

$$R = kN$$

$$N = 6.022 \times 10^{23}$$

$$K = ^\circ\text{C} + 273$$

$$c = 3.0 \times 10^8 \text{ m/s}$$

$$h = 6.626 \times 10^{-34} \text{ Js}$$

$$h_{\text{bar}} = h/2\pi = 1.054 \times 10^{-34} \text{ Js}$$

$$m_e = 9.109 \times 10^{-31} \text{ kg}$$

$$m_p = 1.674 \times 10^{-27} \text{ kg}$$

$$R \text{ (Rydberg)} = 3.289 \times 10^{15} \text{ Hz}$$

STP = 1 atm and 273K

$$\text{Water spec. heat} = 4.18 \text{ J/g } ^\circ\text{C}$$

Useful equations:

$$v = c/\lambda$$

$$\Delta E = h v$$

$$1/\lambda = (R/c) (1/n_1^2 - 1/n_2^2)$$

$$E = (hR)(1/n_1^2 - 1/n_2^2)$$

$$v = R(1/n_1^2 - 1/n_2^2)$$

$$T\lambda_{\text{max}} = c_2/5$$

$$0.5mv^2 = hv - \Phi$$

$$\lambda = h/mv = h/p$$

$$\Delta p \Delta x \geq 1/2h\text{bar}$$

$$m \Delta v \Delta x \geq 1/2h\text{bar}$$

$$H\Psi = E\Psi$$

$$E_n = m^2 h^2 / 8mL^2$$

$$k = PV$$

$$V = kT$$

$$P_1 V_1 = P_2 V_2$$

$$V_1/T_1 = V_2/T_2$$

$$P_1 V_1/T_1 = P_2 V_2/T_2$$

$$PV = nRT$$

$$n = g/MW$$

$$\rho = g/ml$$

$$M = n/V$$

$$E = 0.5 mv^2$$

$$v = (3RT/MW)^{0.5}$$

$$(P - n^2 a/V^2)(V - nb) = nRT$$

$$\text{relative rate} = (MW_1/MW_2)^{0.5}$$

$$m_1 v_1^2 = m_2 v_2^2$$

$$\Delta E = \Delta U = q = mC\Delta T \quad \text{at constant volume}$$

$$E = 0.5 kT \quad \text{per degree of freedom}$$

$$\Delta E = \Delta U = q + w = q - P\Delta V$$

$$\Delta H = q = mC\Delta T \quad \text{at constant pressure}$$

$$\Delta H_{\text{rxn}}^\circ = \sum n \Delta H_f^\circ \text{ products} - \sum n \Delta H_f^\circ \text{ reactants}$$

$$\Delta H_{\text{rxn}} = \sum \text{B.E.}_{\text{rxn}} = \sum \text{B.E.}_{\text{react}} - \sum \text{B.E.}_{\text{prod}}$$

$$C = q/\Delta T$$

$$H = E + PV$$

$$\Delta H = \Delta E + \Delta nRT$$

$$w = -P\Delta V = -\Delta n_g RT$$

$$w = -nRT \ln(V_2/V_1)$$

$$\Delta S_{\text{rxn}}^\circ = \sum n S_m^\circ \text{ products} - \sum n S_m^\circ \text{ reactants}$$

$$S = k \ln W$$

$$\Delta S_{\text{sys}} = q_{\text{sys}}/T$$

$$\Delta S_{\text{surr}} = -\Delta H_{\text{sys}}/T$$

$$\Delta S_{\text{univ}} = \Delta S_{\text{sys}} + \Delta S_{\text{surr}}$$

$$\Delta G_{\text{sys}} = \Delta H_{\text{sys}} - T\Delta S_{\text{sys}} \quad (\text{constant } T, P)$$

$$\Delta G_{\text{rxn}}^\circ = \sum n \Delta G_f^\circ \text{ products} - \sum n \Delta G_f^\circ \text{ reactants}$$